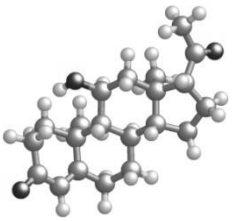


Chem!stry

Name: …… **( )**

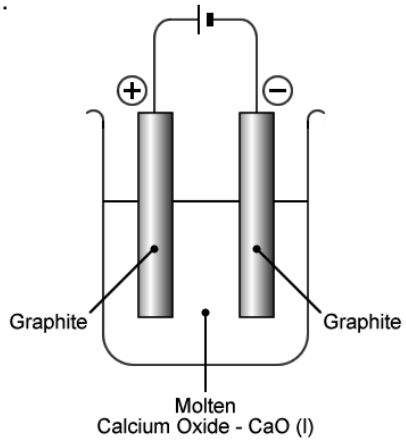
Class: ……

Date: …… **/** …… **/** ……

Revision – Electrolysis of Molten Binary Salts

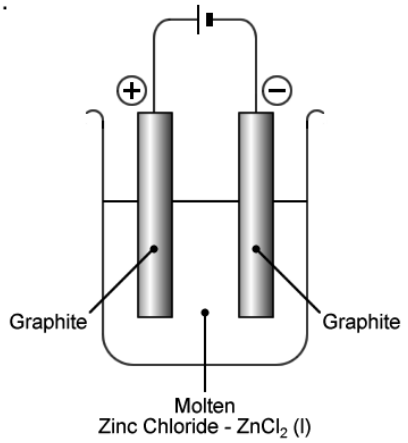
• For the electrolysis of each molten salt; **a)** write the ionic half-equation for the reaction at the anode, **b)** state what you would observe at the anode, **c)** write the ionic half-equation for the reaction at the cathode, **d)** state what you would observe at the cathode.

1.



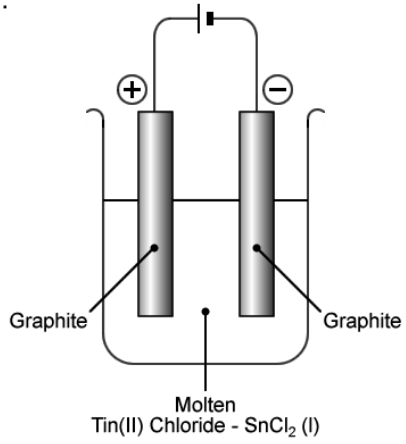
a) Anode (+) ……
b) Observation ……
c) Cathode (–) ……
d) Observation ……

2.



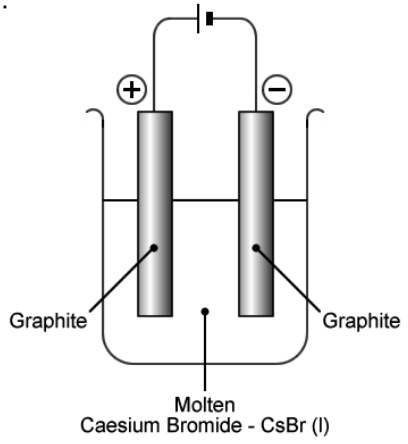
a) Anode (+) ……
b) Observation ……
c) Cathode (–) ……
d) Observation ……

3.



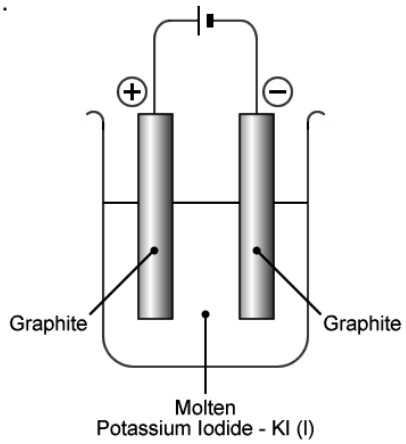
a) Anode (+) ……
b) Observation ……
c) Cathode (–) ……
d) Observation ……

4.



a) Anode (+) ……
b) Observation ……
c) Cathode (–) ……
d) Observation ……

5.



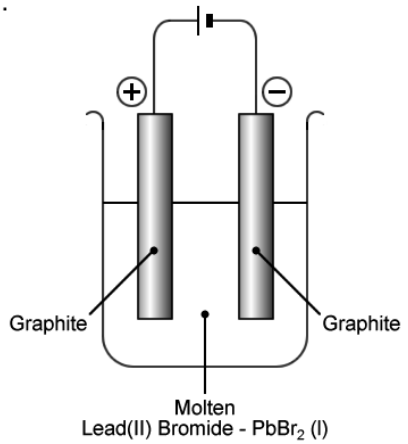
a) Anode (+) .....

b) Observation .....

c) Cathode (–) .....

d) Observation .....

6.



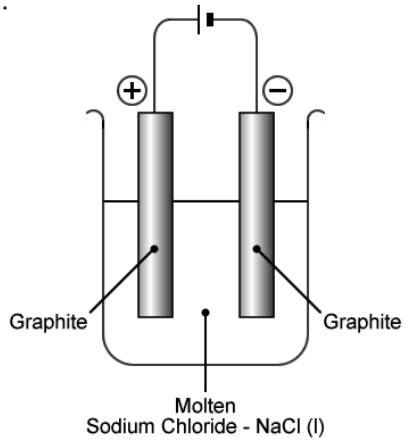
a) Anode (+) .....

b) Observation .....

c) Cathode (–) .....

d) Observation .....

7.



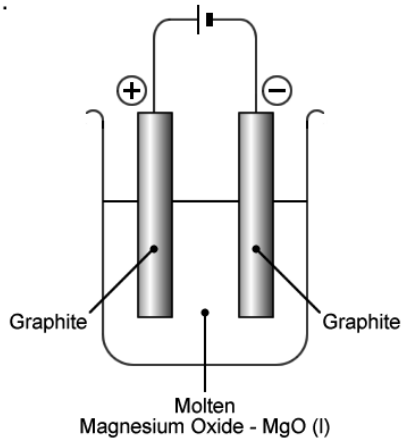
a) Anode (+) .....

b) Observation .....

c) Cathode (–) .....

d) Observation .....

8.



a) Anode (+) .....

b) Observation .....

c) Cathode (–) .....

d) Observation .....

- Scan the QR code given below to view the answers to this assignment.

http://www.chemist.sg/electro_chem/electrolysis_molten_salts_ans.pdf