



Chem!stry

Name: ()

Class:

Date: / /

Assignment on Acids, Bases and Salts #4 – Revision of Key Ideas

1. Define the term *acid*.

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.....

2. Give the names and formulae of some common acids.

- Name: Formula:
- Name: Formula:
- Name: Formula:
- Name: Formula:
- Name: Formula:

3. Define the term *strong acid* and provide two examples to support your answer.

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4. Define the term *weak acid* and provide two examples to support your answer.

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5. Define the term *dibasic acid* and provide one example to support your answer.

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6. a) Give at least one example of a *monobasic acid*.

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b) Give at least one example of a *tribasic acid*.

.....

7. a) Define the term *base*. Give two examples to illustrate your answer.

.....

b) What is an *alkali*? Give two examples to illustrate your answer.

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8. The pH scale.

Acidic						Neutral	Alkaline							
Strongly Acidic			Weakly Acidic				Weakly Alkaline				Strongly Alkaline			
1	2	3	4	5	6	7	8	9	10	11	12	13	14	

What are the approximate pH values of the following solutions?

- a) Aqueous sodium hydroxide
- b) Aqueous sodium chloride
- c) Aqueous ammonia
- d) Rain water
- e) Nitric acid
- f) Ethanoic Acid

9. Classify the following oxides as either *acidic*, *basic*, *neutral* or *amphoteric*.

SO ₃	Na ₂ O	Al ₂ O ₃	SO ₂	PbO
H ₂ O	MgO	CaO	ZnO	NO
CO	P ₄ O ₁₀	CuO	CO ₂	SiO ₂

- Acidic:
- Basic:
- Neutral:
- Amphoteric:

10. Write a balanced chemical equation, including state symbols, for each one of the following chemical reactions:

- a) nitric acid + magnesium → +
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- Which metals do *not* react with acids?
- b) phosphoric acid + sodium hydroxide → +
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- c) ethanoic acid + copper(II) carbonate → + +
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- d) ammonium sulfate + calcium hydroxide → + +
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- Scan the QR code given below for the answers to this assignment.



http://www.chemist.sg/acids/acids_assignments/acids_assignment_4_ans.pdf